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Chapter 12 Stoichiometry 127 SECTION
12.1 THE ARITHMETIC OF EQUATIONS

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(pages 353-358) This section explains how to calculate the amount of reactants required or product formed in a nonchemical process. It teaches you how to interpret chemical equations in terms of interacting moles, representative particles, masses, and gas volume at STP.

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Chapter 3 Stoichiometry: Calculations
with Chemical Formulas and Equations
3.1 Multiple -Choice Questions 1) When
the following equation is balanced, the
coefficients are _____. $C_8H_{18} + O_2 \rightarrow CO$

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2 + H₂O A) 2, 3, 4, 4 B) 1, 4, 8, 9 C) 2,
12, 8, 9 D) 4, 4, 32, 36 E) 2, 25, 16, 18
Answer: E Diff: 2 Page Ref: Sec. 3.1

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Chapter 12 Multiple Choice Identify the letter of the choice that best completes the statement or answers the question.

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____ 1. The first step in most stoichiometry problems is to ____.

- a. add the coefficients of the reagents
- c. convert given quantities to volumes
- b. convert given quantities to moles
- d. convert given quantities to masses

____ 2.

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formula units KClO_3 decom-
pose to form two formula units
 KCl and three molecules O_2
b Four molecules NH_3 react with ...
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