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Assessment Criteria

Unit Atomic Structure

Ib Expectations Atomic Mass Units (AMU):

1/12th of the mass of a carbon - 12 atom in its ground state. This is used to express masses of atomic particles. 1 AMU =

$$1.6605402 \times 10^{-27} \text{ kg}$$

2.1 The nuclear atom -
IB Alchemy Completion
of the Group 4

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Assessment Criteria

Atomic Mass Units
(AMU): 1/12th of the
mass of a carbon - 12
atom in its ground
state. This is used to
express masses of
atomic particles. 1 AMU
 $= 1.6605402 \times 10^{-27}$
kg

**2.1 The nuclear
atom - IB Alchemy**

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LPS Assessment

Unit 1: Atomic
Structure and Nuclear
Unit 2: The Periodic
Table and Electrons
Unit 3: Bonding
Quarter 1 Review Unit
4: Reactions Unit 5:
Chemical Quantities
Unit 6: KMT and GASES
Unit 7:
Thermochemistry Unit
8: Solutions Quarter 3
test Unit 9: Rates and
Equilibrium Unit 10:
Acids and Bases

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Semester 2 Test
Assessment

**Unit 1: Atomic
Structure and
Nuclear - WHS
BLENDED
CHEMISTRY**

IB Chemistry higher level revision notes on atomic theory. Hence, if we know the charge on the nucleus we can use the First Ionization Energies to come to conclusions regarding the distance of the electron from the

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nucleus and possible
interelectronic
repulsion effects..

Hydrogen can be used
as our baseline, as it
has only one proton in
the nucleus and one
electron.

IB Chemistry higher level revision notes: Atomic theory

Completion of the
Group 4 Project and
the IA to be submitted
electronically to IB. If a
year 2 student does

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not complete the Internal Assessment (IA) for a class tested in May 2016, they will be removed from the external exam in May. As such, the student will lose the 1.0 GPA for the class and if applicable, the weight for the first year of the class will also be removed.

Expectations & the IA - Edison IB Physics

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The unified atomic mass unit; The unified atomic mass unit (μ) is commonly used in nuclear physics. It is defined as one twelfth of the mass of a carbon-12 atom. Mass defect and nuclear binding energy; Mass defect. The difference between the mass of an atom and the sum of mass of its constituent parts is called its mass defect.

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**Topic 7: Atomic,
nuclear and particle
physics - IB Physics**

Topic 2: Atomic structure. Notes for the Core IB Chemistry module: Topic 2: Atomic structure. These have been made according to the specification and cover all the relevant topics for examination in May/June.

Topic 2: Atomic structure | A*

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Chemistry

Binding energy: Mass defect (δ): the difference between the total mass of the individual nucleons making up a nucleus and the actual mass of the nucleons. $\delta = \text{total mass of nucleons} - \text{mass of nucleus}$. If it is an atom, remember to subtract the mass of electrons. Einstein's theory of relativity (relationship between mass and energy): $E =$

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Criteria

mc^2 . Binding energy
definition: "the
minimum energy ...

Topic 7: Atomic, Nuclear, and Particle Physics | ib-physics

DP IB CHEM I. When
you miss class it is
your responsibility to
contact me or find out
what you missed. I
expect all students to
keep up with notes and
... UNIT 4: Atomic
Bonding and Structure.

Unit 4 Learning

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lb Expectations
Statements. Unit 4
Notes. VSEPR Shapes.
Atomic Bonding &
Criteria
Structure PPT. UNIT 5:
Energetics and
Thermochemistry .

DP IB CHEMISTRY I - Ms. Peace's Chemistry Class

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LPS

atomic mass unit: The
standard unit that is
used for indicating

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mass of an atom.

History of Dalton's

Atomic Theory

Although the concept of the atom dates back to the ideas of Democritus, the English meteorologist and chemist John Dalton formulated the first modern description of it as the fundamental building block of chemical structures.

History of Atomic

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Structure |

Chemistry [Master]

IB Course Companion

Temp & Thermometers

- Ch 3: pp. 92-95 Heat

& Internal Energy - Ch

3 Kinetic Model - Ch 3:

pp. 106 - 112 ...

Scientists from the

17th and 18th

centuries were working

without the knowledge

of atomic structure and

sometimes developed

theories that were later

found to be incorrect,

such as phlogiston and

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perpetual motion ...

**Unit 12 -
Thermodynamics -
Edison IB Physics**

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Chemistry HL Year 1
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will be a year packed
full of some of the
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and helpful
PowerPoints.

Criteria

IB Chemistry Year 1 - Google Sites

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Assessment Criteria DP
Group 4:Chemistry, DP
- Age 16-18, Objectives
vaporization,
ionization,
acceleration, deflection
and detection. How the
mass spectrometer can
be used to determine
relative

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Ionization:

Vaporization

Ionization

Acceleration ...

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Structure & Isotopes
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IB Chemistry on Atomic
Structure, Particle
Physics and Relative
Atomic Mass 1. Atomic
Structure Unit

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conversion Atomic Size
radius • Order of
magnitude - (10^{-10} -
 10^{-12})m • Radius Li
atom - (1.5×10^{-10})m
• Radius nucleus - ($1 \times$
 10^{-14})m Radius Li
atom Radius Nucleus Li
atom Nucleon - made
up of (protons +
neutrons) Protons -
made up of 2 up
quarks + 1 down quark
Neutron - made ...

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